

## The Ultimate Chemical Equations Handbook Answers Chapter 3

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## **The Ultimate Chemical Equations Handbook Answers Chapter 7**

Note: Not all of the reactions will occur. For those that do not, write no reaction. 1. A piece of copper is dropped into a container of water. No reaction 2. Liquid bromine is added to a container of

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sodium iodide crystals.  $\text{Br (l)} + 2\text{NaI(s)} \rightarrow 2\text{NaBr(s)} + \text{I}_2\text{(s)}$  3. An aluminum strip is immersed in a solution of silver nitrate.

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Re-typed from The Ultimate Chemical Equations Handbook by Hague and Smith 7. Free elements, no matter how complex the molecule, have an oxidation number (valence or charge) equal to zero. OJ. Fe. T~ He < /1 1 ~~ L-td. 8. The following are diatomic or polyatomic elements in nature which must be committed to memory.

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## Quiz on all Chemical Reactions - Max Study

The Ultimate Chemical Equations Handbook SnCO tin(II) carbonate sodium hydrogen carbonate NaHCO 1. 2. 3. 4. 5. 6. 8. 9. 10. 11. vanadium(V) oxide H<sub>2</sub>O dihydrogen monoxide (NH) CO ammonium oxalate 4 224 polonium(VI) thiocyanate Po(SCN)<sub>6</sub> tetraphosphorus decaoxide P<sub>4</sub>O<sub>10</sub> 12. 13. 15. 16. 17. 18. 19. manganese(VII) oxide copper(II) dihydrogen phosphate

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## **KEY to 5 Reaction Prediction - Ipscience.com**

The Ultimate Chemical Equations Handbook Determine the oxidation number of the underlined element:  $\text{KMnO}_4$  Since K is an alkali metal, its charge must be  $1+$  Oxygen is  $2-$  but there are four of them, therefore, 4 times  $2-$  equals  $8-$  If  $1+$  and  $8-$  are added Ultimate Chemical Equation Answers

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The Ultimate Chemical Equations Handbook Determine the oxidation number of the underlined element:  $\text{KMnO}_4$  Since K is an alkali metal, its charge must be 1+ Oxygen is 2— but there are four of them, therefore, 4 times 2— equals 8— If 1+ and 8— are added The Ultimate Chemical Equations Handbook Answers Chapter 10

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